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Solution: $p_{abs} = p_{atm} + p_{gage} = 101.3kPa + 20 \text{ in. H}_2\text{O} \times 1.248 \times 10^{-4} \text{ Pa/in. H}_2\text{O} = 110.26kPa$ Ideal gas law: $\rho = \frac{p}{RT} = \frac{110.26kPa}{1 \text{ kgK} \times 2077 \text{ J} \times 293.2K} = 1.41 \text{ kg/m}^3$ $\rho = 0.181 \text{ kg/m}^3$ b.) Situation: A sphere of 93 mm diameter contains an ideal gas. $T = 20 \text{ C} = 293.2K$ Find: Calculate the density of argon at a vacuum pressure of 8.8 psi. Properties:

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SOLUTION $p=0$. The fluid is considered compressible. $dp = \rho EV$ = However, $V = dv > v$. m . Then, $r \cdot z = h \cdot z - (m > r^2) dp$. $dv = V$. $dr = -m > r$. Therefore, $EV = p \cdot dp > r$ (a) At the surface, where $p \dots$

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